# CHEMISTRY 12th

ASSIGNMENTS OF 'SOLUTION' CHAPTER
part-1

1.Define solution and write their components.
2.Define binary solution.
3.what is homogeneous mixture and heterogeneous mixture?
4.Define Concentration term and what's the way through which we can describe the concentration of the solution quantitatively.
5.write the different types of solution with examples.
6.what is mole fraction? Explain with derivation
7.Define Molarity and molality. Explain ,how the morality is different from molality?
8.calculate the morality of a solution containing 50g of a NaOH In 500ml of solution.
9.Calculate (a) molarity (b)molality (c) mole fraction,of Kl if the density of 20percent(mass/mass) aqueous kl is 1.202 gm/ml

**# CHEMISTRY Xllth. #**
**# part2# (solution)**
**# Assignment #**

1.Define solubility.
2.states Henry's law and mention some important application?
3.why do gases always tend to less Soluble in liquids as the temperature is raised ?
4.why naphthalene and anthracene ball dissolve in Benzene but sodium chloride do not?
5.To increase the solubility of CO2 in Soft drinks and soda water,what is use in bottle?
6.Define
a, Anoxia b, Bends.
7.write the Raoults law of partial vapour pressure of a solution containing the non-volatile solute.
6.write about the ideal solution and non-ideal solution with examples.
7.what is azeotropic solution and give the example of nitric acid?
8.Define colligative properties and write the types of all colligative Properties.
9.Explain the term
a, Relative lowering of vapour pressure.
b, Depression of freezing point.
Or elevation of boiling point.
10.what is osmotic pressure?

**#  CHEMISTRY XIIth  #**

Part-3

**# assignment of solution #**

1.Define osmosis and osmotic pressure.
2.what is osmotic pressure?
3.what is isotonic solutions and hypertonic solutions?
4.define reverse osmosis and explain, how the sea water is purified?
5.explain ,Abnormal molar mass And van't of factor.